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Research Article

## Crystallinity Improvement of $\text{Co}_3\text{O}_4$ by Adding Thiourea

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### ABSTRACT

Tricobalt tetraoxide ( $\text{Co}_3\text{O}_4$ ) samples having different thiourea/Co molar ratio of 0, 5 and 10 were prepared by wet chemical synthesis. The effects of thiourea content on the crystal structure-related parameters of  $\text{Co}_3\text{O}_4$  were determined. The increase in the amount of thiourea caused a gradual decrease in the lattice parameters and specific surface area and an increase in the crystallinity and crystallite size. The experimental analysis results showed that thiourea content can be used to control the crystal structure-related parameters of  $\text{Co}_3\text{O}_4$ .

**Keywords:** Crystal structure, Electron microscopy, X-ray techniques

## Tiyöüre İlavesiyle $\text{Co}_3\text{O}_4$ 'ün Kristalleşmesinin Geliştirilmesi

### ÖZET

Tiyöüre/Co molar oranı 0, 5 ve 10 olan trikobalt tetraoksit ( $\text{Co}_3\text{O}_4$ ) numuneleri yaş kimyasal sentez ile hazırlandı. Tiyöüre içeriğinin  $\text{Co}_3\text{O}_4$ 'ün kristal yapısıyla ilgili parametreleri üzerine etkileri belirlendi. Tiyöüre miktarındaki artış, örgü parametreleri ve spesifik yüzey alanında kademeli bir düşüşe, kristalleşme ve kristal büyüklüğünde bir düşüşe neden oldu. Deneysel analiz sonuçları, tiyöüre içeriğinin  $\text{Co}_3\text{O}_4$ 'ün kristal yapısıyla ilgili parametrelerinin kontrol edilmesinde kullanılabileceğini gösterdi.

**Anahtar Kelimeler:** Kristal yapı, Elektron mikroskopisi, X-ışını teknikleri

## **I. INTRODUCTION**

Cobalt oxide ( $\text{Co}_3\text{O}_4$ ), which is a metal oxide semiconductor (p-type) having a normal spinel structure ( $\text{AB}_2\text{O}_4$  type of  $\text{CoCo}_2\text{O}_4$ ) has got a lot of tremendous properties such as high catalytic reactivity, fast redox behavior, high electrochemical stability, low cost and environmentally friendly and has been used as a gas sensing material owing to its abundant oxygen adsorption and multivalence properties, an anode material for lithium-ion batteries and an electrode for high performance supercapacitors and a catalyst for the oxygen evolution reaction [1-6].  $\text{Co}_3\text{O}_4$  has been prepared via several methods including polyol, hydrothermal, sol-gel, combustion, microemulsion, chemical vapor deposition, sol-flame, solvothermal and sonochemical [7,8].

In the present paper, we aimed to prepare the  $\text{Co}_3\text{O}_4$  powders via a wet chemical method using the precursor of thiourea, which has been used as a sulfur-contained inexpensive organic fuel [9], of various amounts, and we reported the effect of the addition of thiourea on the crystal structure-related parameters of  $\text{Co}_3\text{O}_4$ .

## **II. MATERIALS AND METHOD**

Three  $\text{Co}_3\text{O}_4$  samples with various thiourea ( $\text{CS}(\text{NH}_2)_2$ , Sigma-Aldrich,  $\geq 99.0\%$ ) /cobalt (II) nitrate hexahydrate ( $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ , Sigma-Aldrich,  $\geq 99.0\%$ ) molar ratio of 0, 5 and 10 were produced by wet chemical synthesis, and these samples were referred to as TU0, TU5, and TU10, respectively. To prepare the thiourea-free  $\text{Co}_3\text{O}_4$  sample, a solution of 50 mL of 0.05 M  $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$  in the distilled water was prepared and stirred at  $100^\circ\text{C}$  for 1 h. Then, this mixture was dried in an oven at  $150^\circ\text{C}$  for 15 h and calcined in an electric furnace at  $750^\circ\text{C}$  for 2 h. In this way, the thiourea-free  $\text{Co}_3\text{O}_4$  was obtained. In the synthesis processes of the thiourea-containing  $\text{Co}_3\text{O}_4$  samples, the appropriate amounts of the solutions of 0.25 M and 0.50 M  $\text{CS}(\text{NH}_2)_2$  were added to  $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$  solutions. The above process was repeated to have the thiourea-containing  $\text{Co}_3\text{O}_4$  samples.

X-ray diffraction (XRD) data of the as-synthesized  $\text{Co}_3\text{O}_4$  samples were collected by a Rigaku RadB-DMAX II model diffractometer using  $\text{CuK}\alpha$  radiation. Fourier transform infrared (FTIR) data were recorded by a PerkinElmer Spectrum One spectrophotometer in the mid-infrared spectral range using the KBr pellet technique. A LEO EVO 40xVP scanning electron microscope (SEM) equipped with a Röntech xflash energy dispersive X-ray (EDX) analyzer was used to investigate the morphology and chemical composition of the as-observed areas.

## **III. RESULTS AND DISCUSSION**

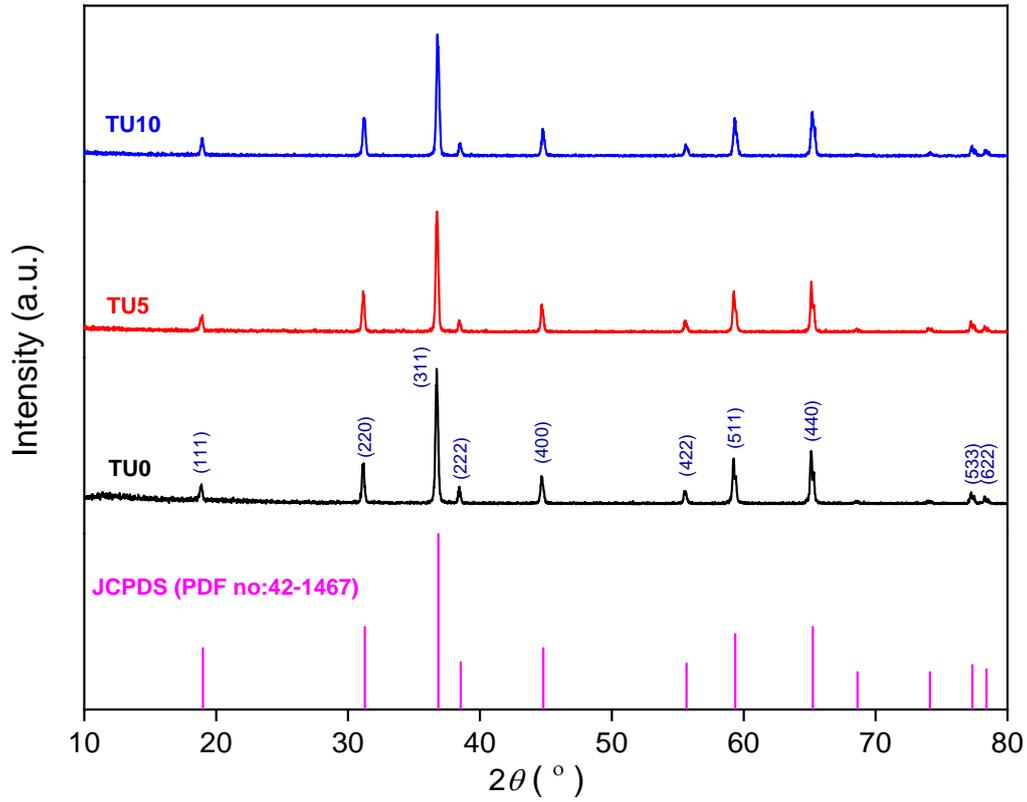
### **A. XRD RESULTS**

The XRD patterns shown in Fig. 1 represent the polycrystalline structure for each sample, and all the as-observed peaks on these patterns are perfectly matched with the reported pattern belonging to the standard pattern of  $\text{Co}_3\text{O}_4$  (JCPDS PDF No: 42-1467).

The estimation of the crystallinity percent ( $X_C\%$ ) can be found in elsewhere [10]. The crystallite size was calculated by using the Williamson-Hall equation as  $D$  [10]:

$$\beta \cos \theta = \frac{0.9\lambda}{D} + 4\varepsilon \sin \theta \quad (1)$$

where  $\beta$ ,  $\lambda$ ,  $\theta$  and  $\varepsilon$  are the full width at half maximum, X-ray wavelength, Bragg angle and lattice strain, respectively. The  $\varepsilon$  and  $D$  values were estimated from the  $\beta\cos\theta$  vs.  $4\sin\theta$  plot given in Fig. 2.



**Figure 1.** XRD patterns of the as-produced samples

The dislocation density ( $\delta$ ) was computed by using the following relation [11]

$$\delta = \frac{1}{D^2} \quad (2)$$

The lattice parameter ( $a$ ), unit cell volume ( $V$ ) and X-ray density ( $\rho_x$ ) of the cubic crystal system were calculated using the following relations, respectively [12,13]:

$$a = d\sqrt{h^2 + k^2 + l^2} \quad (3)$$

$$V = a^3 \quad (4)$$

$$\rho_x = \frac{8M}{N_A a^3} \quad (5)$$

where  $h$ ,  $k$  and  $l$  are the Miller's indices,  $M$  is the molecular weight and  $N_A$  is the Avogadro's constant. The specific surface area ( $SSA$ ) for each sample was calculated using the following relation [13]:

$$SSA = \frac{6}{D\rho_x} \quad (6)$$

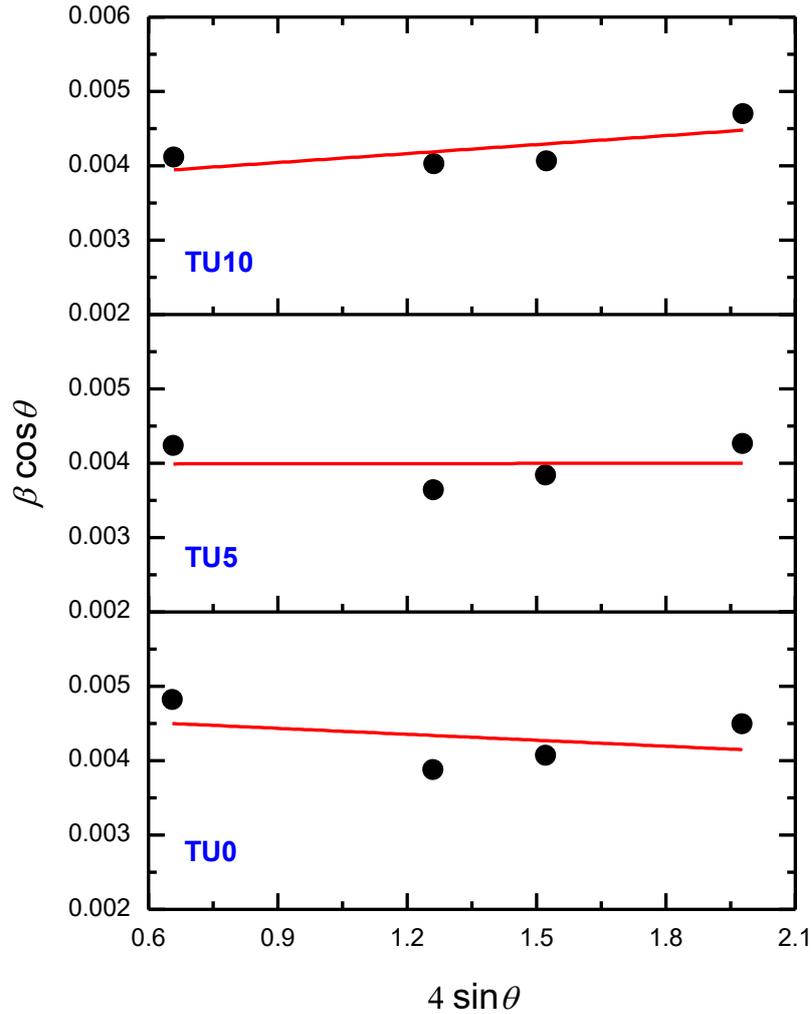


Figure 2. Williamson-Hall plot of  $\text{Co}_3\text{O}_4$  samples

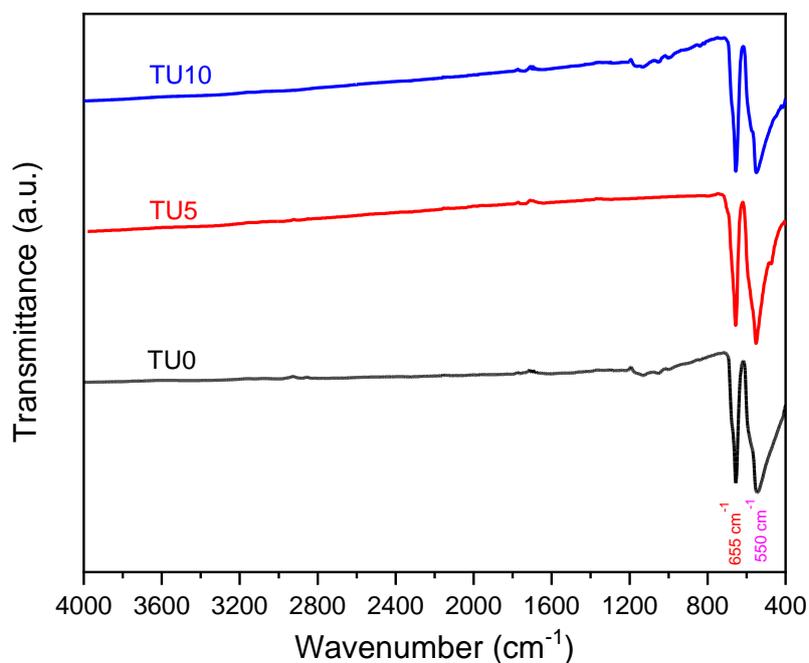
Table 1. The calculated values of the XRD related parameters of  $\text{Co}_3\text{O}_4$  samples

Sample	$D$ (nm)	$\epsilon$	$\delta$ ( $\text{m}^{-2}$ )	$X_C\%$	$a$ (nm)	$V$ ( $\text{nm}^3$ )	$\rho_X$ ( $\text{kg m}^{-3}$ )	$SSA$ ( $\text{m}^2 \text{kg}^{-1}$ )
TU0	29.63	$-2.67 \times 10^{-4}$	$11.39 \times 10^{14}$	92.1	0.8111	0.5336	5995	33,778
TU5	34.75	$8.07 \times 10^{-4}$	$8.28 \times 10^{14}$	93.6	0.8106	0.5326	6006	28,748
TU10	36.67	$4.02 \times 10^{-4}$	$7.44 \times 10^{14}$	94.0	0.8098	0.5310	6024	27,162

The as-calculated values of all the above-mentioned parameters are given in Table 1. The  $X_C\%$ ,  $D$  and  $\rho_X$  increase with an increasing amount of thiourea, while the  $a$ ,  $V$  and  $SSA$  decrease. These findings are in a very good harmony with the results belonging to the fuels of aspartic acid, glycine and sucrose [14-16]. The changes reported in Table 1 verify that the fuel to oxidizer ratio effects all the as-calculated parameters. This result is in a good agreement with Venkateswara Rao and Sunandana [17]. Carvalho et al. [18] reported that the excessive amount of the combustion fuel of urea improves the crystallinity. As can be seen from Table 1, our results are in a good agreement with their report. The as-obtained negative lattice strain for the sample TU0 indicates that the unit cell under compressive stress, and its positive value means that the unit cell under the tensile stress [19].

## B. FTIR ANALYSIS

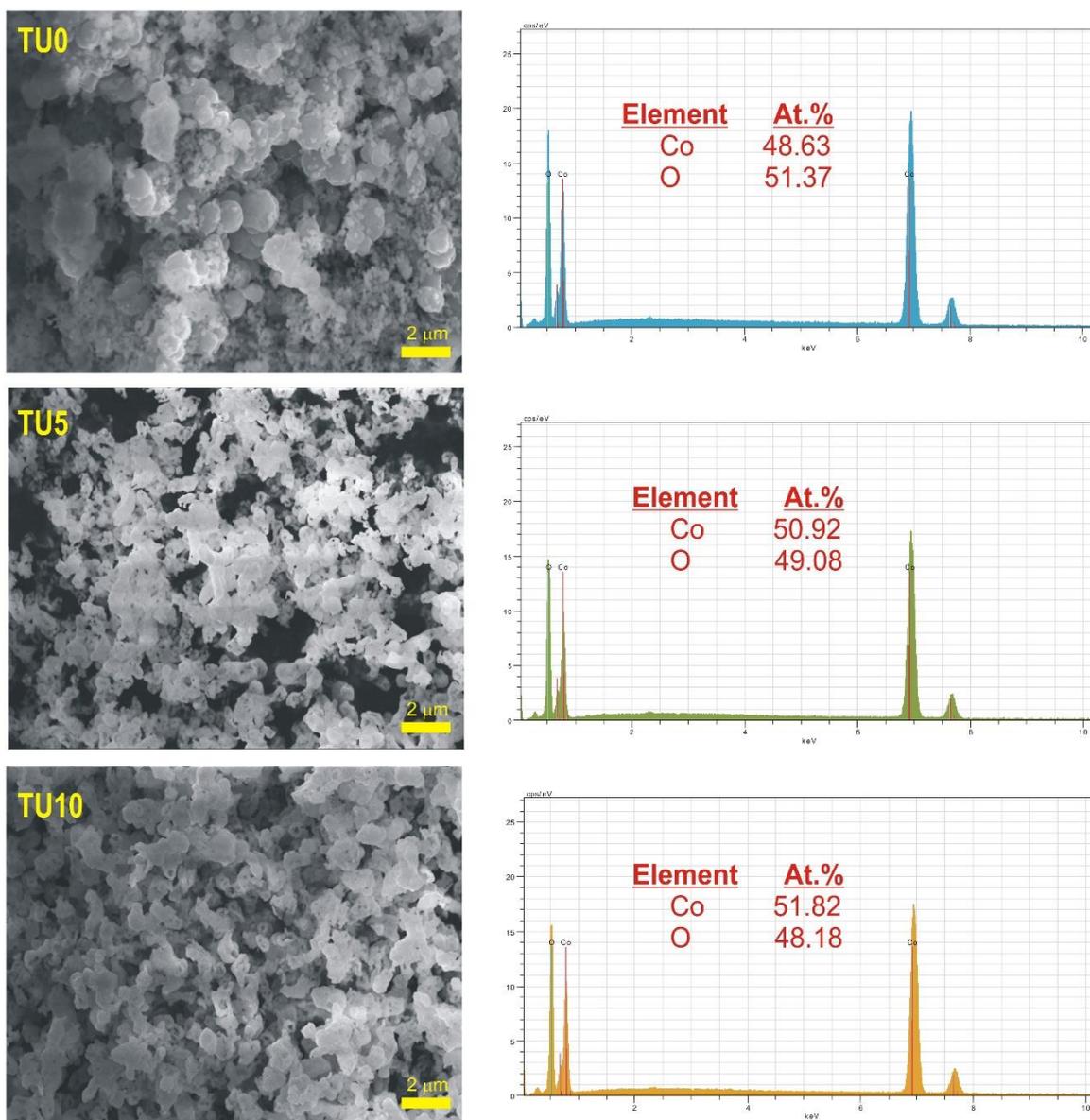
The FTIR spectra for each  $\text{Co}_3\text{O}_4$  sample are shown in Fig. 3. Two sharp bands, which are related to the vibration modes occurred between metal and oxygen bonds [20], are observed in the as-mentioned spectra for each sample. The bands indicating the formation of  $\text{Co}_3\text{O}_4$  structure are detected at  $655\text{ cm}^{-1}$  and  $550\text{ cm}^{-1}$ , and these are assigned to O-Co-O bridging vibration and Co-O stretching vibrations, respectively [21,22].



*Figure 3. The recorded FTIR spectra of  $\text{Co}_3\text{O}_4$  samples*

## C. MORPHOLOGICAL OBSERVATIONS

The SEM images of the samples shown in Fig. 4 point out that all the samples have fine-sized particle distribution, and the thiourea content affects the morphology. The particle size distributions are found to be in the ranges of 85–1,638 nm for TU0, 48–1,830 nm for TU5 and 55–1,990 nm for TU10, respectively. The as-observed particle size distributions are in an agreement with Bazrafshan et al. [23]. The EDX data support that all the samples are composed of Co and O and no impurity is detected.



**Figure 4.** Morphological observation and elemental analysis results of  $\text{Co}_3\text{O}_4$  powders with various thiourea contents

## **IV. CONCLUSIONS**

In this study,  $\text{Co}_3\text{O}_4$  powders with a different amount of thiourea were easily prepared. After analyzing the experimental data, it can be reached the following results. The crystallite size, crystallinity, lattice parameter, unit cell volume and X-ray density increase with the increase in the thiourea content, whereas the specific surface area and dislocation density decrease. The XRD and FTIR results confirm the formation of  $\text{Co}_3\text{O}_4$  without any impurity. The morphology is affected by the thiourea content. The grain size distributions are found to be in the range of about 50-2,000 nm. It is found that the structural properties of  $\text{Co}_3\text{O}_4$  can be improved and controlled by the amount of thiourea.

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